

Science | Worksheet | Class X
Periodic Classification of Elements

1. Fill in the blanks in the following statements:

- The horizontal rows in a periodic table are called _____.
- In going across a period (right to left) in periodic table, the atomic size of the atom _____.
- On moving from right to left in the second period, the number of valence electrons _____.
- Ongoing down in a group in the periodic table, the metallic character of elements _____.
- The tendency to gain an electron _____ on moving down in a group of the periodic table.

2. The atomic numbers of the three elements X, Y and Z are 2, 6 and 10 respectively.

- Which two elements belong to the same group?
- Which two elements belong to the same period?

3. An element X belongs to 3rd period and group 2 of the periodic table. State:

- number of valence electrons
- valency
- metal or non-metal
- name of the element

4. An element Y is in second period and group 16 of the periodic table:

- Is it a metal or non-metal?
- What is the number of valence electrons in its atom?
- What is its valency?
- What is the name of the element?
- What will be the formula of the compound formed by Y with sodium?

5. In each of the following pairs, choose the atom having the bigger size:

- Mg (At. No.12) or Cl (At. No. 17)
- Na (At. No. 11) or K (At. No. 19)

6. What is the usual number of valence electrons and valency of group 18 elements of the periodic table?

7. What happens to the number of valence electrons in the atoms of elements as we go down in a group of the periodic table?

8. Name the element which is in:

- First group and third period.
- Seventeenth group and second period.