

Chemical Reactions & Equations Science | Worksheet | Class X

1. Given the following equation:

Zinc + Hydrobromic acid \rightarrow Zinc bromide + Hydrogen

- Write the equation using symbols and formula.
- Name a solid reactant in the reaction.
- Name a gaseous product formed in the reaction.

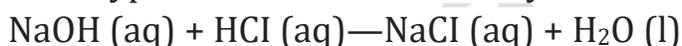
2. Balance the following equations after writing them as molecular equations.

- Sodium + Nitrogen \rightarrow Sodium nitride
- Aluminum + Chlorine \rightarrow Aluminum chloride
- Iron (III) oxide (ferric oxide) + Hydrogen \rightarrow Iron + Water
- Phosphorous + oxygen \rightarrow Phosphorus pentoxide
- Trilead tetraoxide \rightarrow Lead monoxide + Oxygen.

3. Define the following terms

- Thermal decomposition
- Electrolytic decomposition

6. What type of reaction is this? Why is it called so?



4. Give the balanced equation for reactions that take place when the following are heated

- Silver oxide
- Potassium nitrate
- Lead nitrate
- Trilead tetraoxide
- Copper carbonate

5. What information is conveyed by a chemical equation?

6. What information does the following equation convey?



7. A student prepares aqueous solutions of the following salts: Copper sulphate, ferrous sulphate, Sodium sulphate, barium chloride. Write the colour of each solution thus formed.

8. A solution of a substance 'X' is used for white washing.

- Name the substance 'X' and write its formula.
- Write the reaction of the substance 'X' named in (i) above with water.

8. Why is the amount of gas collected in one of the test tubes in Activity 1.7(NCERT text book) double of the amount collected in the other? Name this gas.

9. True and False statement

- a. Chemical changes are generally accompanied by an exchange of energy.
- b. A physical change is usually reversible.
- c. When a magnesium ribbon burns brilliantly in air a physical change takes place.
- d. Ammonium chloride dissolves in water with the absorption of heat. This is an exothermic.

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