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حل التمرين 03

$$M(H_2O) = M(O) + 2M(H) = 18 \text{ g mol}^{-1} \quad .1$$

$$n(H_2O) = \frac{m}{M} \Rightarrow n(H_2O) = \frac{1,8}{18} = 0,1 \text{ mol} \quad .2$$

$$n(H_2O) = \frac{m}{M} \Rightarrow n(H_2O) = \frac{1,5 \cdot 10^3}{18} = 83,34 \text{ mol} \quad .3$$

.4

$$n(H_2O) = \frac{m}{M} \Rightarrow m = M \cdot n(H_2O)$$

$$\Rightarrow m = 18 \times 3,5 \cdot 10^{-2} \text{ g} \Rightarrow \boxed{m = 0,63 \text{ g}}$$

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